



To the Student:

After your registration is complete and your proctor has been approved, you may take the Credit by Examination for IPC 1A.

WHAT TO BRING

- lined notebook paper
- sharpened No. 2 pencils
- a **non-programmable** calculator

ABOUT THE EXAM

The examination for the first semester of Integrated Physics and Chemistry 1A consists of 85-91 multiple choice questions, which will make up 75% of the final exam grade; the other 25% will be skills and essay questions. You will be provided with a Scantron® answer sheet for the objective questions; you will answer the skills and essay questions on your own paper. Skills questions will cover such things as measuring, calculating, graphing, designing an experiment, etc.

The exam is based on the Texas Essential Knowledge and Skills (TEKS) for this subject. The full list of TEKS is included in this document (it is also available online at the Texas Education Agency website, <http://www.tea.state.tx.us/>). The TEKS outline specific topics covered in the exam, as well as more general areas of knowledge and levels of critical thinking. Use the TEKS to focus your study in preparation for the exam.

The examination will take place under supervision, and the recommended time limit is three hours. You may not use any notes or books. You may use a **non-programmable** calculator. A percentage score from the examination will be reported to the official at your school.

In preparation for the examination, review the TEKS for this subject. All TEKS are assessed. It is important to prepare adequately. Since questions are not taken from any one course, you can prepare by reviewing any of the state-adopted textbooks that are used at your school. The textbook used with our IPC 1A course is:

McLaughlin, Charles William, et al. (2005). *Physical Science*. New York: Glencoe McGraw Hill. ISBN 9780078600517.

In order to be successful on the examination, you must study and review all of the concepts of Integrated Physics and Chemistry. These concepts are listed on pages that follow this letter.

Good luck on your examination!

IPC 1A CBE—Study Sheet

Before taking this exam, you should be able to:

- plan and implement an investigation using scientific method, formulate hypotheses, select appropriate equipment and technology, draw inferences, and communicate valid conclusions;
- make wise choices in the use and conservation of resources, disposal or recycling of materials (especially as they apply to substances studied, such as petrochemicals and polymers);
- demonstrate safe practices for lab investigations;
- collect data and make measurements with precision;
- identify tools to measure length, volume, mass and temperature;
- identify metric units and explain the importance of a universal system of measurement;
- use dimensional analysis to convert units;
- organize, analyze, evaluate, make inferences and predict trends from data (using skills such as reading data charts, tables, and the periodic table);
- define matter, describe properties of each class of matter, list the general properties of matter, describe the relationship between mass and weight (gravity), and calculate density;
- identify and describe four phases of matter, phase changes, and energy changes that take place during phase changes;
- define chemical and physical changes, and recognize and identify those changes;
- list the contributions of major scientists such as Archimedes, Rutherford, Thomson, Bohr, the Curies, Einstein, Dalton, and Becquerel;
- describe the historical development of the atomic theory; analyze, critique, review, and compare atomic theories; and explain how indirect evidence has led to understanding of atom;
- identify three subatomic particles, relate charges of each, relative sizes, placement in atom, and role in bonding, nuclear fusion, and nuclear fission;
- identify properties of fluids, such as density, viscosity, and buoyancy;

- classify substances as elements and compounds or mixtures (heterogeneous or homogeneous);
- describe the general properties of elements and relate the chemical behavior of elements, including bonding (ionic, covalent, or metallic), to their placement on the periodic table;
- analyze energy changes that accompany chemical changes as exothermic or endothermic;
- identify or be able to draw atoms, ions, and isotopes;
- define radioactivity;
- identify and describe the law of conservation of mass;
- describe types of nuclear reactions such as fission and fusion and know some of their applications in medicine, energy production, tracing, etc.
- relate the structure of water to its role as universal solvent;
- evaluate the environmental and economic impact of end products of chemical reactions;
- recognize and understand reading of chemical equations, understand the function of coefficients and subscripts, recognize balanced chemical equations, and identify the type of chemical reaction as synthesis, decomposition, single replacement, or double replacement;
- relate concentration of ions in a solution to physical and chemical properties such as pH;
- state the properties of acids and bases, have a general knowledge of the placement of some common acids and bases on the pH scale, and relate acid rain to the pH scale;
- recognize how various factors affect solubility (including temperature, pressure, particle size, nature of solute, and solvent);
- explain how petroleum products are separated from crude oil, explain what petrochemicals and polymers are, and describe the economic and environmental impact of oil spills and of disposing or recycling such materials.

PERIODIC TABLE OF THE ELEMENTS

PERIODIC TABLE OF THE ELEMENTS

GROUP	GROUP NUMBERS																18																			
	IUPAC RECOMMENDATION (1985)																VIIIA																			
1	1A																2																			
1	1	1.0079															2	4.0026																		
1	H																He																			
	HYDROGEN																HELIUM																			
2	3	6.941	4	9.0122															9	18.998	10	20.180														
	Li		Be																F		Ne															
	LITHIUM		BERYLLIUM																FLUORINE		NEON															
3	11	22.990	12	24.305															17	35.453	18	39.948														
	Na		Mg																Cl		Ar															
	SODIUM		MAGNESIUM																CHLORINE		ARGON															
4	19	39.098	20	40.078	21	44.956	22	47.867	23	50.942	24	51.996	25	54.938	26	55.845	27	58.933	28	58.933	29	63.546	30	65.39	31	69.723	32	72.64	33	74.922	34	78.96	35	79.904	36	83.80
	K		Ca		Sc		Ti		V		Cr		Mn		Fe		Co		Ni		Cu		Zn		Ga		Ge		As		Se		Br		Kr	
	POTASSIUM		CALCIUM		SCANDIUM		TITANIUM		VANADIUM		CHROMIUM		MANGANESE		IRON		COBALT		NICKEL		COPPER		ZINC		GALLIUM		GERMANIUM		ARSENIC		SELENIUM		BROMINE		KRYPTON	
5	37	85.468	38	87.62	39	88.906	40	91.224	41	92.906	42	95.94	43	(98)	44	101.07	45	102.91	46	106.42	47	107.87	48	112.41	49	114.82	50	118.71	51	121.76	52	127.60	53	126.90	54	131.29
	Rb		Sr		Y		Zr		Nb		Mo		Tc		Ru		Rh		Pd		Ag		Cd		In		Sn		Sb		Te		I		Xe	
	RUBIDIUM		STRONTIUM		YTRIUM		ZIRCONIUM		NIOBIUM		MOLYBDENUM		TECHNETIUM		RUTHENIUM		RHODIUM		PALLADIUM		SILVER		CADMIUM		INDIUM		TIN		ANTIMONY		TELLURIUM		IODINE		XENON	
6	55	132.91	56	137.33	57-71	72	178.49	73	180.95	74	183.84	75	186.21	76	190.23	77	192.22	78	195.08	79	196.97	80	200.59	81	204.38	82	207.2	83	208.98	84	(209)	85	(210)	86	(222)	
	Cs		Ba		La-Lu		Hf		Ta		W		Re		Os		Ir		Pt		Au		Hg		Tl		Pb		Bi		Po		At		Rn	
	CAESIUM		BARIUM		LANTHANIDE		HAFNIUM		TANTALUM		TUNGSTEN		RHENIUM		OSMIUM		IRIDIUM		PLATINUM		GOLD		MERCURY		THALLIUM		LEAD		BISMUTH		POLONIUM		ASTATINE		RADON	
7	87	(223)	88	(226)	89-103	104	(261)	105	(262)	106	(266)	107	(264)	108	(277)	109	(268)	110	(281)	111	(272)	112	(285)													
	Fr		Ra		Ac-Lr		Rf		Db		Sg		Bh		Hs		Mt		Uun		Uuu		Uub													
	FRANCIUM		RADIUM		ACTINIDE		RUTHERFORDIUM		DUBNIUM		SEABORGIUM		BOHRNIUM		HASSIUM		MEITNERIUM		UNUNUNIUM		UNUNBIUM															
LANTHANIDE																																				
6	57	138.91	58	140.12	59	140.91	60	144.24	61	(145)	62	150.36	63	151.96	64	157.25	65	158.93	66	162.50	67	164.93	68	167.26	69	168.93	70	173.04	71	174.97						
	La		Ce		Pr		Nd		Pm		Sm		Eu		Gd		Tb		Dy		Ho		Er		Tm		Yb		Lu							
	LANTHANUM		CERIUM		PRASEODYMIUM		NEODYMIUM		PROMETHIUM		SAMARIUM		EUROPIUM		GADOLINIUM		TERBIUM		DYSPROSIUM		HOLMIUM		ERBIUM		THULIUM		YTTERBIUM		LUTETIUM							
ACTINIDE																																				
7	89	(227)	90	232.04	91	231.04	92	238.03	93	(237)	94	(244)	95	(243)	96	(247)	97	(247)	98	(251)	99	(252)	100	(257)	101	(258)	102	(259)	103	(262)						
	Ac		Th		Pa		U		Np		Pu		Am		Cm		Bk		Cf		Es		Fm		Md		No		Lr							
	ACTINIUM		THORIUM		PROTACTINIUM		URANIUM		NEPTUNIUM		PLUTONIUM		AMERICIUM		CURIUM		BERKELIUM		CALIFORNIUM		EINSTEINIUM		FERMIUM		MENDELEVIUM		NOBELIUM		LAWRENCIUM							

Texas Essential Knowledge and Skills

IPC 1A – Integrated Physics and Chemistry, First Semester

TTU: IPC 1A, CBE, v.3.0		
TEKS: §113.38. Integrated Physics and Chemistry, Beginning with School Year 2010-2011 (One Half Credit)		
TEKS Requirement (Secondary)	Set A Question Numbers	Set B Question Numbers
§112.38. Integrated Physics and Chemistry, Beginning with School Year 2010-2011.		
(a) General requirements. Students shall be awarded one credit for successful completion of this course. Prerequisites: none. This course is recommended for students in Grade 9 or 10.		
(b) Introduction.		
(1) Integrated Physics and Chemistry. In Integrated Physics and Chemistry, students conduct laboratory and field investigations, use scientific methods during investigation, and make informed decisions using critical thinking and scientific problem solving. This course integrates the disciplines of physics and chemistry in the following topics: force, motion, energy, and matter.		
(2) Nature of science. Science, as defined by the National Academy of Sciences, is the "use of evidence to construct testable explanations and predictions of natural phenomena, as well as the knowledge generated through this process." This vast body of changing and increasing knowledge is described by physical, mathematical, and conceptual models. Students should know that some questions are outside the realm of science because they deal with phenomena that are not scientifically testable.		
(3) Scientific inquiry. Scientific inquiry is the planned and deliberate investigation of the natural world. Scientific methods of investigation are experimental, descriptive, or comparative. The method chosen should be appropriate to the question being asked.		
(4) Science and social ethics. Scientific decision making is a way of answering questions about the natural world. Students should be able to distinguish between scientific decision-making methods (scientific methods) and ethical and social decisions that involve science (the application of scientific information).		
(5) Science, systems, and models. A system is a collection of cycles, structures, and processes that interact. All systems have basic properties that can be described in space, time, energy, and matter. Change and constancy occur in systems as patterns and can be observed, measured, and modeled. These patterns help to make predictions that can be scientifically tested. Students should analyze a system in terms of its components and how these components relate to each other, to the whole, and to the external environment.		
(c) Knowledge and skills.		
(1) Scientific processes. The student, for at least 40% of instructional time, conducts laboratory and field investigations using safe, environmentally appropriate, and ethical practices. The student is expected to:		
(A) demonstrate safe practices during laboratory and field investigations; and	1,2,5,69,76, 84	1, 69, 92, 93
(B) demonstrate an understanding of the use and conservation of resources and the proper disposal or recycling of materials.	2,38,	2, 38, 66,
(2) Scientific processes. The student uses scientific methods during laboratory and field investigations. The student is expected to:		
(A) know the definition of science and understand that it has limitations, as specified in subsection (b)(2) of this section;	32,76, 77	76. 77
(B) plan and implement investigative procedures, including asking questions, formulating testable hypotheses, and selecting equipment and technology;	6, 33, 86, 87	78, 79
(C) collect data and make measurements with precision;	3, 4, 6, 30, 31, 67, 68, 70, 75, 86, 87	3, 4, 5, 20, 67, 68, 70, 75, 76, 80
(D) organize, analyze, evaluate, make inferences, and predict trends from data; and	6, 33, 86, 87	79, 92,
(E) communicate valid conclusions.	4, 6, 30, 31, 67, 68, 70, 75, 86, 87	6, 7, 92
(3) Scientific processes. The student uses critical thinking, scientific reasoning, and problem solving to make informed decisions. The student is expected to:		

(A) in all fields of science, analyze, evaluate, and critique scientific explanations by using empirical evidence, logical reasoning, and experimental and observational testing, including examining all sides of scientific evidence of those scientific explanations, so as to encourage critical thinking by the student;	32, 34, 86, 87	6, 33, 34
(B) communicate and apply scientific information extracted from various sources such as current events, news reports, published journal articles, and marketing materials;	82, 83, 84	81, 82, 83,
(C) draw inferences based on data related to promotional materials for products and services;	33, 81, 85	79, 80, 84
(D) evaluate the impact of research on scientific thought, society, and the environment;	74, 85	74, 84
(E) describe connections between physics and chemistry and future careers; and	56, 57,	23, 57
(F) research and describe the history of physics and chemistry and contributions of scientists.	10, 46, 48, 54, 57,	10, 48, 49, 54,
(4) Science concepts. The student knows concepts of force and motion evident in everyday life. The student is expected to:		
(A) describe and calculate an object's motion in terms of position, displacement, speed, and acceleration;		
(B) measure and graph distance and speed as a function of time using moving toys;		
(C) investigate how an object's motion changes only when a net force is applied, including activities and equipment such as toy cars, vehicle restraints, sports activities, and classroom objects;		
(D) assess the relationship between force, mass, and acceleration, noting the relationship is independent of the nature of the force, using equipment such as dynamic carts, moving toys, vehicles, and falling objects;		
(E) apply the concept of conservation of momentum using action and reaction forces such as students on skateboards;		
(F) describe the gravitational attraction between objects of different masses at different distances, including satellites; and		
(G) examine electrical force as a universal force between any two charged objects and compare the relative strength of the electrical force and gravitational force.		
(5) Science concepts. The student recognizes multiple forms of energy and knows the impact of energy transfer and energy conservation in everyday life. The student is expected to:		
(A) recognize and demonstrate that objects and substances in motion have kinetic energy such as vibration of atoms, water flowing down a stream moving pebbles, and bowling balls knocking down pins;		
(B) demonstrate common forms of potential energy, including gravitational, elastic, and chemical, such as a ball on an inclined plane, springs, and batteries;		
(C) demonstrate that moving electric charges produce magnetic forces and moving magnets produce electric forces;		
(D) investigate the law of conservation of energy;		
(E) investigate and demonstrate the movement of thermal energy through solids, liquids, and gases by convection, conduction, and radiation such as in weather, living, and mechanical systems;		
(F) evaluate the transfer of electrical energy in series and parallel circuits and conductive materials;		
(G) explore the characteristics and behaviors of energy transferred by waves, including acoustic, seismic, light, and waves on water as they superpose on one another, bend around corners, reflect off surfaces, are absorbed by materials, and change direction when entering new materials;		
(H) analyze energy conversions such as those from radiant, nuclear, and geothermal sources; fossil fuels such as coal, gas, oil; and the movement of water or wind; and		
(I) critique the advantages and disadvantages of various energy sources and their impact on society and the environment.		
(6) Science concepts. The student knows that relationships exist between the structure and properties of matter. The student is expected to:		
(A) examine differences in physical properties of solids, liquids, and gases as explained by the arrangement and motion of atoms, ions, or molecules of the substances and the strength of the forces of attraction between those particles;	11, 24, 49, 53, 55,	13, 15, 22, 43, 52

(B) relate chemical properties of substances to the arrangement of their atoms or molecules;	14, 49, 53, 55, 60, 61, 62, 63, 66,	44, 46, 53, 63
(C) analyze physical and chemical properties of elements and compounds such as color, density, viscosity, buoyancy, boiling point, freezing point, conductivity, and reactivity;	6, 7, 8, 9, 50,	8, 9, 50, 51, 65
(D) relate the physical and chemical behavior of an element, including bonding and classification, to its placement on the Periodic Table; and	13, 16, 41, 44, 45, 47, 51, 53, 58, 59, 72, 88	11, 12, 16, 17, 45, 47, 58, 61, 72, 78
(E) relate the structure of water to its function as a solvent and investigate the properties of solutions and factors affecting gas and solid solubility, including nature of solute, temperature, pressure, pH, and concentration.	7, 52, 64, 65	6, 25, 26, 27, 29, 42
(7) Science concepts. The student knows that changes in matter affect everyday life. The student is expected to:		
(A) investigate changes of state as it relates to the arrangement of particles of matter and energy transfer;	15, 37, 39, 42,	14, 85, 86, 88
(B) recognize that chemical changes can occur when substances react to form different substances and that these interactions are largely determined by the valence electrons;	12, 36, 39,	39, 87, 88
(C) demonstrate that mass is conserved when substances undergo chemical change and that the number and kind of atoms are the same in the reactants and products;	36, 37, 39,	85, 87, 88
(D) analyze energy changes that accompany chemical reactions such as those occurring in heat packs, cold packs, and glow sticks and classify them as exothermic or endothermic reactions;	19, 79, 80	19, 89, 90,
(E) describe types of nuclear reactions such as fission and fusion and their roles in applications such as medicine and energy production; and	21, 23, 40,	21, 40
(F) research and describe the environmental and economic impact of the end-products of chemical reactions such as those that may result in acid rain, degradation of water and air quality, and ozone depletion.	28, 78	28, 91
<i>Source: The provisions of this §112.38 adopted to be effective August 4, 2009, 34 TexReg 5063.</i>		